

6CH01/01

Section A

Question Number	Correct Answer	Reject	Mark
1	A		1

Question Number	Correct Answer	Reject	Mark
2	C		1

Question Number	Correct Answer	Reject	Mark
3	C		1

Question Number	Correct Answer	Reject	Mark
4	B		1

Question Number	Correct Answer	Reject	Mark
5	D		1

Question Number	Correct Answer	Reject	Mark
6	B		1

Question Number	Correct Answer	Reject	Mark
7	B		1

Question Number	Correct Answer	Reject	Mark
8	D		1

Question Number	Correct Answer	Reject	Mark
9	A		1

Question Number	Correct Answer	Reject	Mark
10	A		1

Question Number	Correct Answer	Reject	Mark
11	B		1

Question Number	Correct Answer	Reject	Mark
12	A		1

Question Number	Correct Answer	Reject	Mark
13	C		1

Question Number	Correct Answer	Reject	Mark
14	B		1

Question Number	Correct Answer	Reject	Mark
15	A		1

Question Number	Correct Answer	Reject	Mark
16	B		1

Question Number	Correct Answer	Reject	Mark
17	D		1

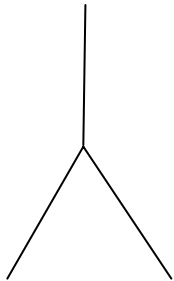
Question Number	Correct Answer	Reject	Mark
18	C		1

Question Number	Correct Answer	Reject	Mark
19	C		1

Question Number	Correct Answer	Reject	Mark
20	D		1

Section B

Question Number	Correct Answer	Reject	Mark
21 (a)(i)	Easier to transport / easier to store / less space / less volume needed for storage / easier to handle / easier to transfer <i>IGNORE</i> references to “safety” Accept Denser/cheaper to transport OWTTE	Just “cost”	1

Question Number	Correct Answer	Reject	Mark
21 (a)(ii)	skeletal formula (1)  Name: butane (1) Stand alone skeletal formula (1)  Name: methylpropane OR 2-methylpropane (1) <i>IGNORE</i> incorrect punctuation [e.g. extra/missing hyphens, etc.] Stand alone <i>IGNORE</i> displayed formulae if also given with skeletal formulae if 2 correct displayed formulae are given max 1 out of 2 for the structures		4

Question Number	Correct Answer	Reject	Mark
21 (a)(iii)	(Structural) isomers		1

Question Number	Correct Answer	Reject	Mark
21 (b)(i)	$\text{Cl}_2 \rightarrow \text{Cl}^\bullet + \text{Cl}^\bullet /$ $\text{Cl}_2 \rightarrow 2\text{Cl}^\bullet \quad (1)$ (U.V.) light / sunlight (1) Must show the dots • <i>IGNORE</i> any subsequent propagation steps in (b)(i)	heat alone	2

Question Number	Correct Answer	Reject	Mark
21 (b)(ii)	$\text{C}_3\text{H}_8 + \text{Cl}^\bullet \rightarrow \text{C}_3\text{H}_7^\bullet + \text{HCl} \quad (1)$ $\text{C}_3\text{H}_7^\bullet + \text{Cl}_2 \rightarrow \text{C}_3\text{H}_7\text{Cl} + \text{Cl}^\bullet \quad (1)$ Must show the dots •		2

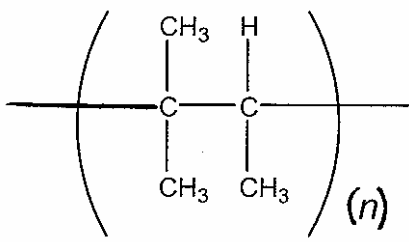
Question Number	Correct Answer	Reject	Mark
21 (b)(iii)	$\text{C}_3\text{H}_7^\bullet + \text{Cl}^\bullet \rightarrow \text{C}_3\text{H}_7\text{Cl}$ OR $\text{Cl}^\bullet + \text{Cl}^\bullet \rightarrow \text{Cl}_2$ OR $\text{C}_3\text{H}_7^\bullet + \text{C}_3\text{H}_7^\bullet \rightarrow \text{C}_6\text{H}_{14}$ Must show dots in termination step		1

Question Number	Correct Answer	Reject	Mark
21 (c)(i)	Alkene / triene Accept Diene Carbon-carbon double bond		1

Question Number	Correct Answer	Reject	Mark
21 (c)(ii)	From: Red / brown / orange / yellow or combinations of these colours To: colourless <i>both colours needed</i>	“clear” instead of colourless	1

Question Number	Correct Answer	Reject	Mark
21 (c)(iii)	Electrophilic (1) addition (1)		2

Question Number	Correct Answer	Reject	Mark
21 (c)(iv)	Calculation: 0.01 mol myrcene reacts with 0.03 mol H₂ OR 1 mol myrcene reacts with 3 mol H₂ (1) Structural formula: (CH₃)₂CH(CH₂)₃CH(CH₃)CH₂CH₃ OR <pre> H H H H H H H H₃C—C—C—C—C—C—C—H CH₃ H H H CH₃ H H </pre> (1) Accept Fully displayed formula/skeletal formula Mark calculation and structural formula independently.		2

Question Number	Correct Answer	Reject	Mark
21 (d)	 repeat unit (1) continuation bonds shown (but these bonds do not have to cut through the brackets) (1) <i>n</i> not essential <i>IGNORE</i> the position of “<i>n</i>” relative to the repeat unit (e.g. can be written as a superscript)		2

Question Number	Correct Answer	Reject	Mark																					
22 (a)(i)	<table><tr><th>Energy change</th><th>Letter</th><th>$\Delta H/\text{kJ mol}^{-1}$</th></tr><tr><td>Lattice energy for sodium chloride</td><td>E</td><td>-775</td></tr><tr><td>Enthalpy change of atomization of sodium</td><td>C</td><td>+109</td></tr><tr><td>Enthalpy change of atomization of chlorine</td><td>A</td><td>+121</td></tr><tr><td>First ionization energy of sodium</td><td>B</td><td>+494</td></tr><tr><td>First electron affinity of chlorine</td><td>F</td><td></td></tr><tr><td>Enthalpy change of formation of sodium chloride</td><td>D</td><td>-411</td></tr></table>	Energy change	Letter	$\Delta H/\text{kJ mol}^{-1}$	Lattice energy for sodium chloride	E	-775	Enthalpy change of atomization of sodium	C	+109	Enthalpy change of atomization of chlorine	A	+121	First ionization energy of sodium	B	+494	First electron affinity of chlorine	F		Enthalpy change of formation of sodium chloride	D	-411		3
	Energy change	Letter	$\Delta H/\text{kJ mol}^{-1}$																					
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	Enthalpy change of atomization of chlorine	A	+121																					
	First ionization energy of sodium	B	+494																					
	First electron affinity of chlorine	F																						
	Enthalpy change of formation of sodium chloride	D	-411																					
	6 correct letters (3)																							
5 or 4 correct letters (2)																								
3 or 2 correct letters (1)																								
1 or 0 correct letters (0)																								

Question Number	Correct Answer	Reject	Mark
22 (a)(ii)	Expression such as: $D = C + B + A + F + E$ $- 411 = + 109 + 494 + 121 + F + (- 775)$ $F = - 411 - 109 - 494 - 121 + 775$ (1) Answer: $F = -360 \text{ (kJ mol}^{-1}\text{)} \text{ (1)}$ <i>Check empty box in 22(a)(i), as answer may be written there.</i> Answer must follow from working Correct answer only (2) Correct answer with some consistent working (2)		2

Question Number	Correct Answer	Reject	Mark
22 (b)(i)	(Bonding in NaCl) 100% ionic OR almost completely ionic OR no covalent character/(very) little covalent character	'Molecule' (0)	1

Question Number	Correct Answer	Reject	Mark
22 (b)(ii) QWC	AgI has (a degree of) covalent character (1) due to polarization or distortion (of the anion) (1)		2

Question Number	Correct Answer	Reject	Mark
22 (c) QWC	Any two of the following: (outermost) electron further from the nucleus/atoms get bigger/more shells (outermost) electron more shielded (by inner shells of e⁻) (force of) attraction between nucleus and (outermost) electron decreases (down the Group) OR (outermost) electron held less strongly (down the Group) OR (outermost) electron becomes easier to remove (down the Group) IGNORE any references to (effective) nuclear charge or more protons.	“ions” get bigger (down Group)	2

Question Number	Correct Answer			Reject	Mark
23 (a)	element	structure	bonding		3
	sodium	Giant	metallic		
	silicon	Giant (atomic)/ macromolecular/ giant molecular	covalent		
	sulfur	simple / small molecules OR (simple) molecular OR S ₈ molecules	covalent or van der Waals' forces/ London forces/ intermolecular forces/dispersion forces/induced- dipole forces		
	IGNORE the word “lattice” OR “crystalline” 6 boxes correct (3) 5,4 boxes correct (2) 3,2 boxes correct (1) 1,0 boxes correct (0)				

Question Number	Correct Answer	Reject	Mark
23 (b)	Si : covalent bonds / many bonds / strong bonds (between atoms) (1) S : weak forces /van der Waals' forces/London forces/dispersion forces/intermolecular forces/induced-dipole forces (1) (need to be overcome)	any reference to intermolecular forces in Si suggestion that covalent bonds are broken	2

Question Number	Correct Answer	Reject	Mark
23 (c) QWC	Cations/ions decrease in size (from Na^+ to Al^{3+}) OR charge increases/charge density on (cat)ions increases/ “effective nuclear charge” increases (from Na^+ to Al^{3+}) (1) more e^- (per atom in ‘sea’ of delocalized electrons) / more delocalized electrons OR (force of) attraction between (cat)ions/nucleus and (delocalised) electrons increases (from Na to Al) (1) IGNORE “nuclear charge increases”/ “increasing no. of protons”	atoms decrease in size any mention of “molecules”/ “covalent bonds”/ “van der Waals’ forces”/ “ionic bonds” (0) overall	2

Question Number	Correct Answer	Reject	Mark
23 (d)(i) QWC	- Add MgO to acid/react MgO with acid/dissolve MgO in acid (1) [NOTE: mention of heating not required. IGNORE water bath/reflux] - Filter (1) Heat/boil filtrate /MgSO₄ solution (until volume reduced by half) (1) - Leave to cool/leave to crystallise/evaporate slowly/leave to evaporate (1) (decant / filter / pick out crystals, then) Leave to dry/pat dry/dry between filter papers/put in an oven/put in a desiccator/dry the crystals (1) IGNORE any washing of crystals immediately prior to drying them	Just “warm” the filtrate/MgSO₄ solution Use of a desiccant (added to crystals)	5

Question Number	Correct Answer	Reject	Mark
23 (d)(ii)	Rinse with (plenty of) water /use a damp cloth or damp (paper) towel / add a (named) weak alkali (e.g. solid or aqueous sodium hydrogencarbonate)	Any named strong alkali/just “strong alkali”	1

Question Number	Correct Answer	Reject	Mark
23 (e)(i)	Insoluble strontium sulfate/insoluble SrSO₄ (forms on the strontium carbonate)		1

Question Number	Correct Answer	Reject	Mark
23 (e)(ii)	$\text{Sr}^{2+}(\text{aq}) + \text{SO}_4^{2-}(\text{aq}) \rightarrow \text{SrSO}_4(\text{s})$ species (1) state symbols (1) 2nd mark is cq on first mark $\text{Sr}^{2+}(\text{aq}) + 2\text{Cl}^{-}(\text{aq}) + 2\text{Na}^{+}(\text{aq}) + \text{SO}_4^{2-}(\text{aq}) \rightarrow \text{SrSO}_4(\text{s}) + 2\text{Cl}^{-}(\text{aq}) + 2\text{Na}^{+}(\text{aq})$ scores (1) $\text{SrCl}_2(\text{aq}) + \text{Na}_2\text{SO}_4(\text{aq}) \rightarrow \text{SrSO}_4(\text{s}) + 2\text{NaCl}(\text{aq})$ scores (1)		2

Question Number	Correct Answer	Reject	Mark
24 (a)(i)	$\frac{2.90}{58} = 0.05(00) \text{ (mol)}$ correct answer only (1)		1

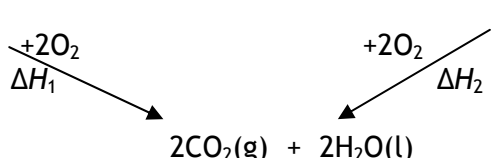
Question Number	Correct Answer	Reject	Mark
24 (a)(ii)	$200 \times 4.18 \times 58.2$ $= 48655 \text{ (J) OR } 48.655 \text{ kJ (1)}$ for correct $\Delta T(1)$ <i>IGNORE</i> sf <i>IGNORE</i> signs at this stage		2

Question Number	Correct Answer	Reject	Mark
24 (a)(iii)	$-\frac{48655}{0.0500} = -973\,100 \text{ (J mol}^{-1}\text{)}$ $= -973 \text{ kJ mol}^{-1} \text{ (3 s.f.)}$ $/ -973000 \text{ J mol}^{-1} \text{ (3 s.f.)}$ answer (1) sign and units (1) [Do not award sign and units mark if units given are just “kJ” or just “J”] three sig figs (1) <i>CQ on (a)(i) & (ii)</i>		3

Question Number	Correct Answer	Reject	Mark
24 (b)(i)	Heat loss/energy loss Accept Incomplete combustion OWTTE <i>IGNORE</i> “experimental error”/ “departure from standard conditions”	Anything related to “average values” (0)	1

Question Number	Correct Answer	Reject	Mark
24 (b)(ii)	Difference: less exothermic / less negative <i>IGNORE</i> “higher” if written with less exothermic/less negative Accept just “lower”/ “less” (1) Justification: energy taken in to form gas/energy required to form gas/energy needed to form gas/takes heat in to form gas/heat required to form gas Or reverse argument (1) <i>Mark these two points independently</i>	Just “higher” (0) Just “H ₂ O(g) is not water’s standard state”	2

Question Number	Correct Answer	Reject	Mark
24 (c)(i)	Enthalpy / energy / heat (energy) change (when) one mole of a substance/one mole of a compound (1) is formed from its elements (in their most stable states) (1) 298K / 25°C / a stated temperature AND 1 atm pressure/100 kPa (1) <i>IGNORE</i> any references to concentration	“energy required” OR “energy released” “one mole of product(s) ” is formed from its reactants room temperature/rtp	3

Question Number	Correct Answer	Reject	Mark
24 (c)(ii)	Cycle or formula expression  $\Delta H_f^\theta = \Delta H_1 - \Delta H_2$ $= (2 \times -394) + (2 \times -286) - (-870)$ $= -490 \text{ (kJ mol}^{-1}\text{)}$ • correct expression or cycle (1) • evidence for doubling both ΔH_c^θ [C] and ΔH_c^θ [H₂] (1) • answer (1) Correct answer with no working scores full marks		3